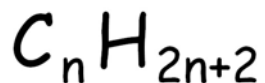


Alkanes

-saturated (all single bonds) $C - C$

General formula:



Naming Alkanes:

IUPAC name	Number of Carbon Atoms	Molecular formula
Methane	1	CH_4
Ethane	2	C_2H_6
Propane	3	C_3H_8
Butane	4	C_4H_{10}
Pentane	5	C_5H_{12}
Hexane	6	C_6H_{14}
Heptane	7	C_7H_{16}
Octane	8	C_8H_{18}
Nonane	9	C_9H_{20}
Decane	10	$C_{10}H_{22}$

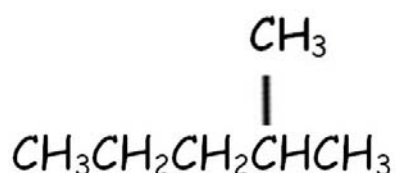
*These names serve as parent names for all non-aromatic compounds

Systemic nomenclature devised by IUPAC for any acyclic compound:

1) Identify the longest un-branched carbon chain. This establishes the parent name (based on the alkane)

2) Everything else on the carbon chain is treated as a substituent

ex.



Parent chain has 5 carbons so the parent name is pentane. A substituent is also present.

How do we name substituents?

Naming Substituents

a) halogens

F fluoro Cl chloro

Br bromo I iodo

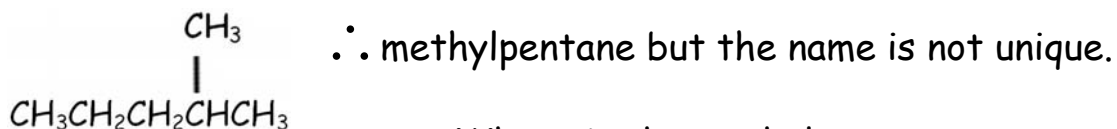
Subtract "ine" from the end of the name and add "o"

b) carbon substituent (alkyl group)

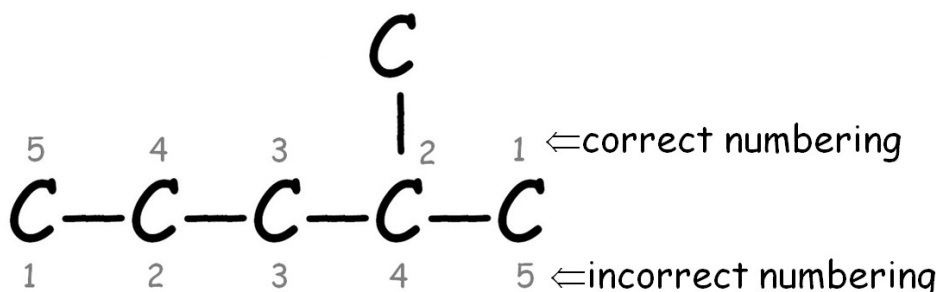
To name a substituent alkyl group you substitute the "ane" from the end of its parent alkane name with "yl"

1 carbon	$-\text{CH}_3$	methyl	
2 carbons	$-\text{CH}_2\text{CH}_3$	ethyl	
3 carbons	$-\text{CH}_2\text{CH}_2\text{CH}_3$	propyl	} two possibilities
	CH_3CHCH_3 	isopropyl	

return to the compound:

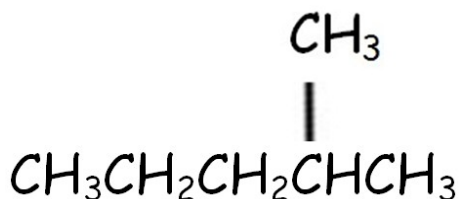


Where is the methyl?



numbering begins at the end nearest the branch

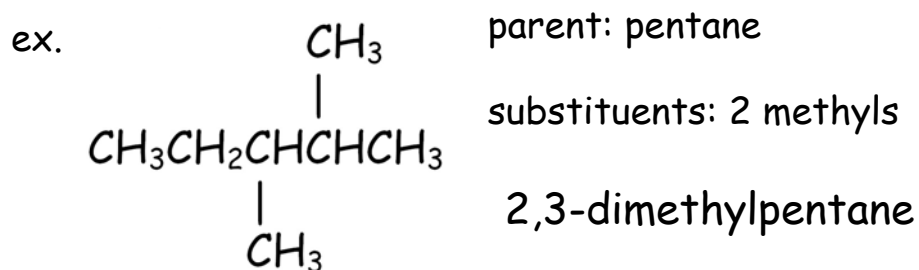
back to the compound



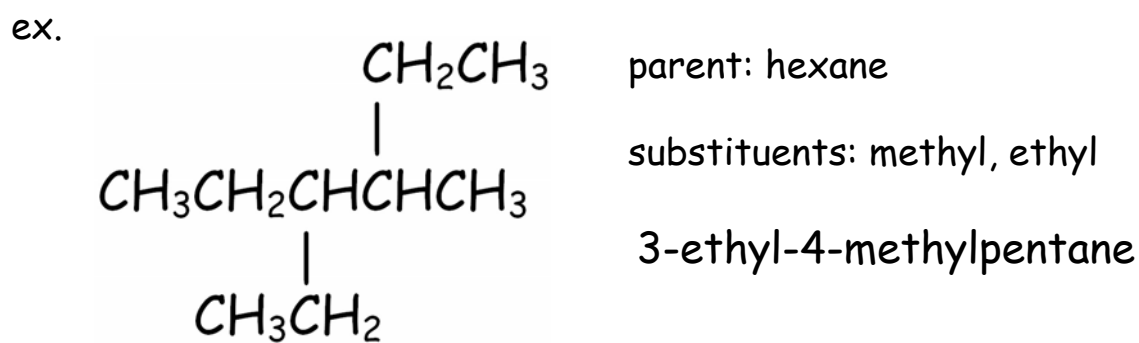
2-methylpentane



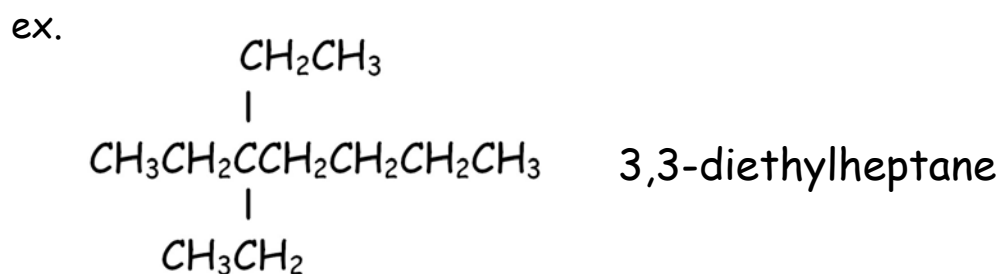
The 2 identifies the carbon of the parent chain that the methyl substituent is attached to.



note: numbers are separated from letters by a hyphen and numbers are separated from each other by a comma



note: when two or more substituents are present the order they appear in the name is determined alphabetically



Organic Nomenclature

note:-separate numbers with commas

-separate numbers and words with hyphens

-a number must appear for every branch

-number the chain so that the lowest set of numbers appear in the name