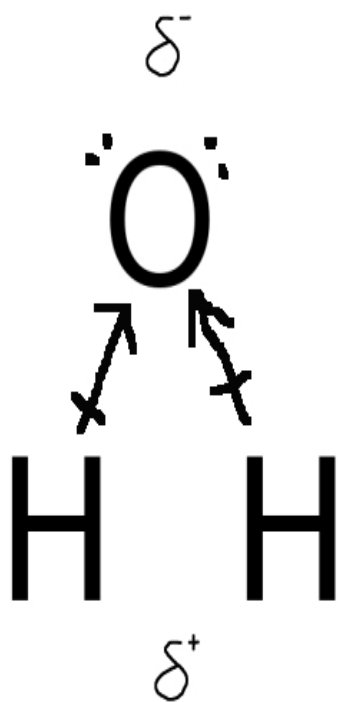


WATER:

THE UNIVERSAL SOLVENT

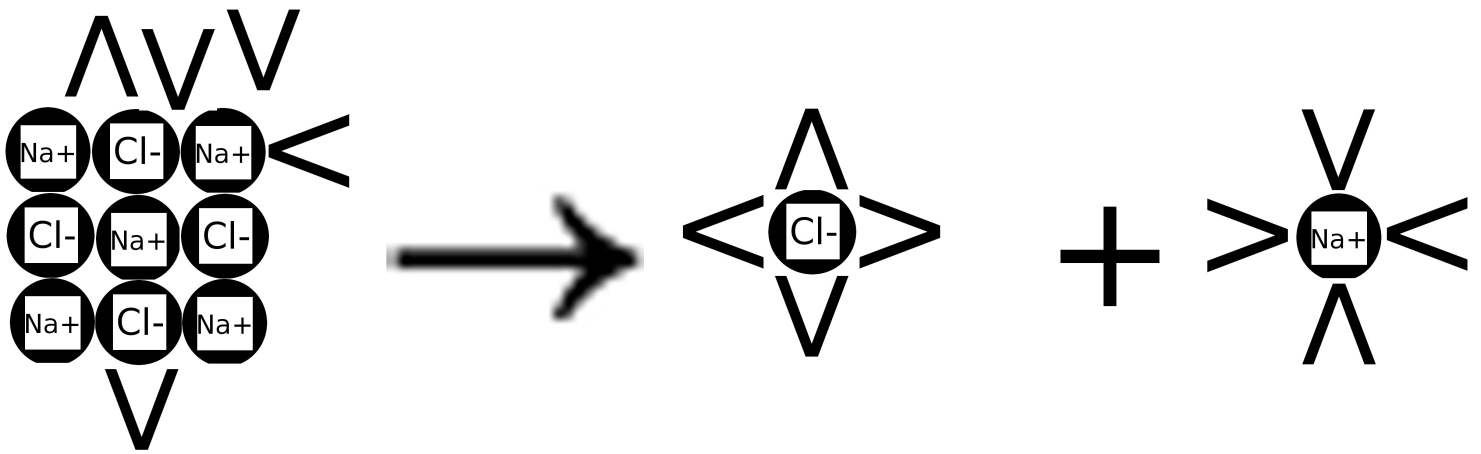
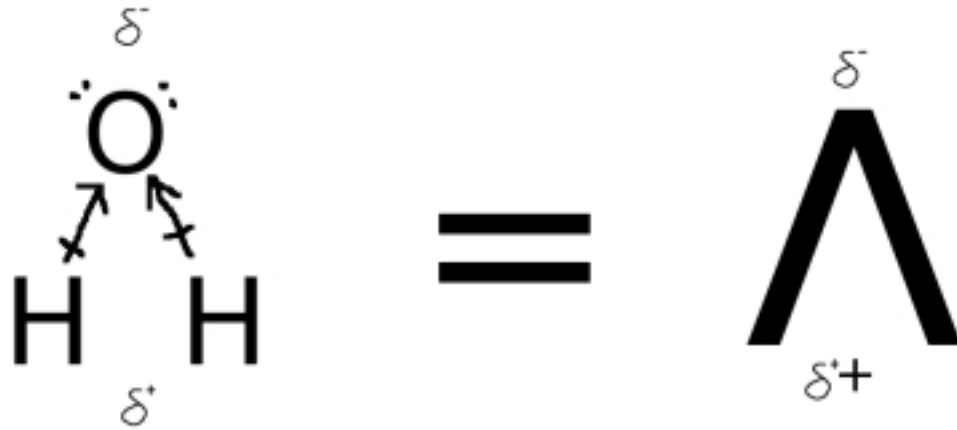
-Water is a bent molecule containing two oxygen-hydrogen bonds



-oxygen has high electronegativity therefore it attracts the two shared pair of electrons more... because of this the oxygen end of the molecule is partially negative (δ^-) and the hydrogen end is partially positive (δ^+).

-Water is a polar molecule, it has a positive end (+) and a negative end(-)

-Water is a polar molecule, it has a positive end (+) and a negative end(-)



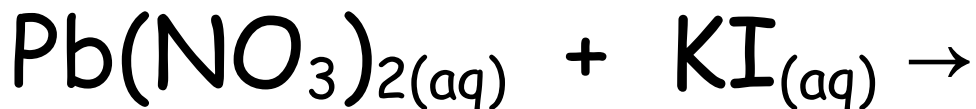
-attraction between the ions and water is able to overcome the attractive forces between the sodium ions and the chlorine ions and the ions surrounded by water move off into solution thus dissociation has occurred

-other ionic substances dissociate in a manner similar to sodium chloride; however some dissolve more readily and some dissolve less readily

PRECIPITATE REACTIONS

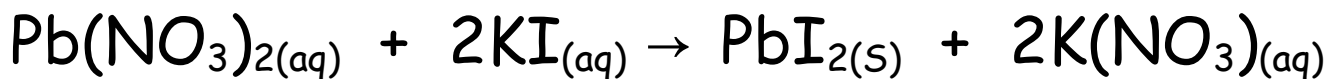
precipitate - a solid that is formed as a result of mixing two solutions

Consider the following reaction:



- 1.) complete the balanced double displacement rxn
- 2.) Identify the precipitate
- 3.) Write the ionic equation
- 4.) Write the net ionic equation
- 5.) Identify the spectator ions

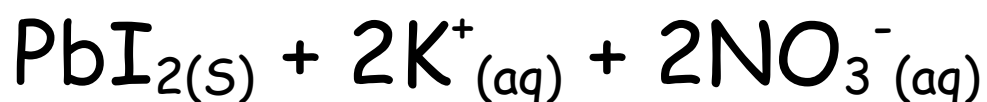
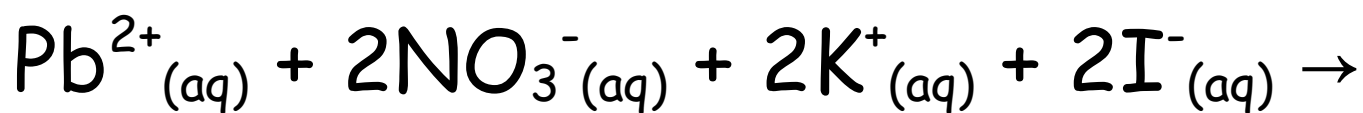
1.) balanced equation (from solubility table):



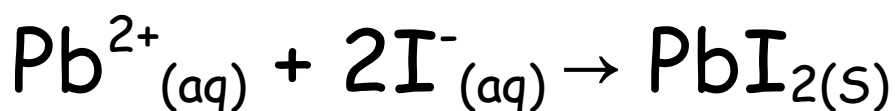
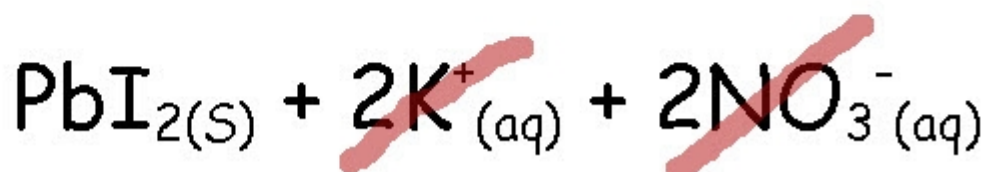
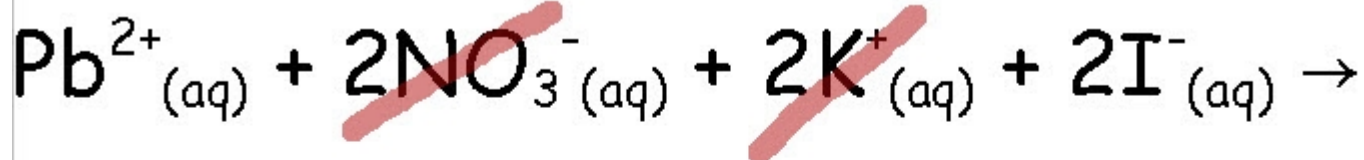
↑

2.) precipitate (insoluble)

3.) ionic equation:



4.) net ionic equation (leave out ions on both sides of the balanced ionic equation):



5.) spectator ions (the ions on both sides of the net ionic equation):

